Note: Although Chem 1306 and 1106 are listed as two separate courses, they are in fact linked together in content in such a way that you <u>must</u> take them simultaneously. You will receive the same grade for Chem 1106 as for Chem 1306. Thus, if you drop one course without dropping the other, you will automatically fail the one you did not drop. The records office will not automatically drop one course if you drop the other. If you are not already enrolled in both Chem 1306 and 1106, contact Dr. Maxwell or your academic advisor to remedy this situation.

Dr. Janet L. Maxwell   janet.maxwell@angelo.edu
Office Hours (on collaborate)
MWF 9:00 – 10:00 am
TR   9:30 – 10:30 am
or by appointment

Required Texts:

 "General, Organic and Biological Chemistry", 4th edition by Frost and Deal
ISBN 978-0-13-498869-6

"Lab Manual for Chem 1306/1106, 5th edition", Maxwell
ISBN 978-0-13-594071-6

Also <u>Required</u>:   1) Registration/Participation in the Achieve Learning For-Credit Homework Software System
2)   Enrollment in Top Hat Classroom Response System
no device needed
3) A scientific calculator   (Texas Instruments is the recommended brand, but other brands are acceptable)
4) Web cam and <u>required software for Respondus Lock-down Browser and monitor.</u>

Course Description:

An Introduction to General, Organic and Biological Chemistry with emphasis on the role of chemistry in health and illness. Intended for nurses and allied health professionals.

Student Learning Outcomes:

A.   Introduction to Chemistry - Students will be able to:
1. Classify a sample of matter as a mixture, a compound, or an element; and describe a mixture as homogenous or heterogenous.
2. Know the names and symbols of common elements, the differences between metals, metalloids and nonmetals, and the organization of the periodic table.
3. Convert a measurement from one unit to another using a conversion factor, report a measured value to the correct number of digits, be able to use a calculator to carry out calculations, be able to round answers to the appropriate number of significant figures, be able to express numbers in scientific notation, to enter numbers in scientific notation correctly into the calculator, to understand the values of numbers expressed using scientific notation and to understand how to find the percent of a whole or the percent that one number is of another number.

4. Define mass and volume, and know how to express these properties using the metric system and define and solve for density and specific gravity and use density or specific gravity to solve for the mass or volume.
5. Understand and interconvert between temperature scales, understand units of energy and be able to calculate the specific heat of a substance.
6. Distinguish the bulk properties of the three states of matter, and relate these properties to the behavior of the particles that make up the substance.
7. Understand the difference between precision and accuracy.
8. Carry out dosage calculations and use percents in health care applications.
9. Determine whether a process is a physical change or a chemical reaction and write and balance a chemical equation to represent a chemical reaction.

B. Atoms and Radioactivity – Students will be able to:

1. Know the properties of the subatomic particles, and determine the mass and charge of an atom from the number of each type of subatomic particle.
2. Understand the concepts of atomic number and mass number and how these relate to the numbers of protons and neutrons in the nucleus of the atom.
3. Know how isotopes are related to one another, and understand the relationship of atomic weight to the masses of individual atoms.
4. Understand the concepts of radioactivity and radioisotopes and be able to describe and write the equations for the common types of nuclear decay.
5. Understand the units of radioactive disintegration and the concept of half-life.
6. Understand the most common medical applications for radioisotopes.

C. Compounds – Students will be able to:

1. Understand how electrons are arranged in an atom, write electron arrangements for the first 20 elements and understand the octet rule.
2. Predict ion charges for representative elements, and understand how atoms gain and lose electrons to form ions.
3. Predict the formulas of ionic compounds using the ion charges, and learn the names and charges of common transition metal ions.
4. Write the names and formulas of ionic compounds that contain two elements, learn the names and formulas of common polyatomic ions, and write the names and formulas of compounds that contain polyatomic ions.
5. Understand how covalent bonds are formed, use the octet rule to predict the number of covalent bonds an atom can form, draw Lewis structures for molecules that contain single bonds, draw Lewis structures for molecules that contain double or triple bonds, use lines to represent bonding electrons and write names for binary covalent compounds.
6. Use the mole concept to express numbers of formula units, interconvert between moles and masses, understand the significance of Avogadro's number and use it to interconvert between numbers of atoms and masses in grams.
7. Determine the shapes of molecular compounds. use electronegativities to predict whether a covalent bond is polar or nonpolar, determine the type of charge on each bonding atom and determine the polarity of molecules containing more than one covalent bond.

D. Introduction to Organic Compounds – Students will be able to:

1. Explain why carbon is uniquely suited to be the main structural element of organic chemistry, and describe the ways in which carbon atoms form covalent bonds.
2. Learn the names of the first 10 linear alkanes, and use common conventions to draw their structural formulas.
3. Classify fatty acids based on their structures.
4. Understand how and why chemists use functional groups to classify organic molecules and how alcohols and amines are categorized as primary, secondary or tertiary.

5. Distinguish linear and branched alkanes and cycloalkanes, and recognize and draw isomers of simple alkanes.
6. Name branched alkanes, cycloalkanes and haloalkanes.
7. Name and draw the *cis* and *trans* forms of cycloalkanes and alkenes, and distinguish between constitutional isomers and *cis-trans* isomers.
8. Understand the concept of chirality in organic chemistry and be able to distinguish stereoisomers from other types of isomers.

E. Carbohydrates – Students will be able to:
1. Understand the structural features and typical physical properties of monosaccharides, including the chirality of these molecules.
2. Understand the concept of organic oxidation and reduction reactions as applied to the concept of reducing sugars.
3. Understand the formation of ring structures in monosaccharides and the relationship between the anomers of a monosaccharide ring.
4. Understand, name and draw the structures of the most common glycosidic linkages in disaccharides and understand the breaking and forming of glycosidic bonds.
5. Describe the building blocks, linkages and biological functions of the common disaccharides and polysaccharides.
6. Understand the structure and function of the carbohydrate antigens in blood.

F. Intermolecular Forces – Students will be able to:
1. Describe the different types of intermolecular forces between molecules or ions, and relate the strength of these forces to physical properties.
2. Name and define the phase changes.
3. Understand the rule of "like dissolves like" and how it applies to nonpolar compounds, polar compounds and ionic compounds, recognize hydrophilic and hydrophobic regions in a molecular compound, and rank the solubilities of structurally related compounds.
4. Understand the role of soap in removing oil and grease.
5. Understand the role of the strength of intermolecular forces in determining the melting and boiling points of substances.
6. Understand and apply the relationships between pressure, volume and temperature for a gas.
7. Understand the qualitative and quantitative aspects of Boyle's Law and Charles' Law and be able to predict the behavior of gases under changing conditions.
8. Understand the effect of saturation and unsaturation of fats on their melting points and distinguish between *cis*-unsaturated and *trans*-unsaturation fats.
9. Understand the structure and function of lipids in cell membranes.

G. Solution Chemistry – Students will be able to:
1. Identify the solute and solvent in a solution and distinguish between solutions, colloids and suspensions.
2. Describe and interpret the solubility of a compounds, and predict the effects of temperature and pressure on solubility.
3. Describe what happens when ionic and molecular compounds dissolve in water and distinguish between strong electrolytes, weak electrolytes and nonelectrolytes.
4. Determine the number of equivalents per mole for dissolved ions.
5. Calculate and use milliequivalents per liter, millimoles per liter, molarity, percent concentrations, other common concentration units involving masses of solutes.
6. Calculate the final volume or concentration of a solution in a dilution, and calculate the volumes of a concentrated solution and water needed to carry out a dilution.
7. Determine the direction of osmosis and dialysis, and predict the effect of a solution on red blood cells using the overall molarity of the solution.
8. Understand and explain the different modes of transports of substances across cell membranes.

Student Learning Outcomes will be evaluated using exam questions.

Grading:

Two One-Hour Exams	2 x 150 = 300 pts
(Each one hour exam will be worth 12 % of the overall grade)	
One Two-Hour Final Exam (comprehensive)	1 x 250 = 250 pts
(The Final Exam will be worth 20 % of the final grade)	
Achieve Learning Online Homework Grade	100 pts
(The Achieve Homework will be worth 8 % of the final grade)	
Top Hat Participation Grade	100 pts
(The Top Hat Participation will be worth 8 % of the final grade)	
Quizzes	10 x 15 = 150 pts
(11 quizzes will be given and the lowest quiz grade will be dropped – the sum of the quiz grades will be worth 12 % of the final grade)	
Lab Grade	14 x 25 = 350 pts
(14 labs will be completed and the one lowest lab grade will be replaced with a perfect score* – the sum of the lab grades will be worth 28 % of the final grade)	_____
	Total = 1250 pts

*Only students with a valid documented excuse can have one lab grade replaced with a perfect score

Students are responsible for monitoring their own grades on Blackboard. The category in Blackboard that accurately reports the current grade is the "Weighted Average". This column does not always appear on cell phones, but it does appear on a computer.

Total averages within the following ranges at the end of the semester guarantee the student at least the indicated letter grade:

A   90% of the total or better
B   80-89% of the total
C   70-79% of the total
D   60-69% of the total
F   less than 60% of the total

One Hour Exams:

The schedule for the one hour exams can be found on the schedule page for your section in this syllabus. The one hour online exams will open at 7:30 am on the day of the test and will close at the official end of class. Each one hour exam will consist of 30 multiple choice questions worth 5 points each. Questions will be in a random, scrambled order.

Final Exam:

See the schedule for the date and time of your final exam. The final exam will consist of 50 questions worth 5 points each. Half the final exam will consists of problems covered in the class after Exam 2. The other half of the exam will consist of problems from the material covered before Exam 2. Questions will be in a random, scrambled order.

Deadline:

Last Day to drop the course: Thursday, April 28

Persons with disabilities which may warrant academic accommodations must contact the Student Life Office, Room 112, University Center, in order to request and to implement academic accommodations.

Quizzes:

You are expected to learn the material in lab before we cover it in class. You are expected to work independently on quizzes. Online quizzes using the lockdown browser and monitor will be given on Sundays as shown in the course schedule on the last page of this syllabus. Quiz topics will include the material from the previous week's lab and the sections associated with those topics from the textbook as well as additional sections as assigned in class. In order to study for the quizzes, you need to read the textbook sections listed for the quiz topic and work all the practice problems and additional problems listed in "Practice Problems and Additional Problems" handout on Blackboard. Quizzes will be worth 15 points. The quizzes will be available from 10 am until 11:59 pm on Sunday on quiz weekends. The time limit for each quiz is either 15 or 20 minutes, depending on Dr. Maxwell's discretion. You must answer the questions in order, and you cannot backtrack. Ten quizzes will be given and the two lowest quiz grades will be dropped. If you miss the deadline for taking a quiz, make-up quizzes or late quizzes will not be given for any reason.

Respondus Lock Down Browser and Monitor

Students are required to use the Respondus lock down browser and monitor for each quiz and exam. Students are required to install and test this software before the first quiz is given before the administration of the first quiz. If you cannot install the Respondus lock down browser and monitor so that it works correctly for your computer, you are required to find a computer in one of the computer labs on campus or the library that has the lock down browser and monitor working effectively in order to take your quiz or exam. Please see the "Respondus Lock Down Browser and Monitor" link under Resourses in the left column of your lecture Blackboard course.

Practice Problems:

On blackboard there is a document labeled "Practice Problems and Additional Problems by Section", which can be found under course materials. At the end of each section, there is a list of practice problems and a list of additional problems. The answers to the odd-numbered practice problems and additional problems are given at the end of each chapter. You will be responsible for knowing what problems are assigned for each section to study for quizzes and exams. This information is given in the document labeled. "Practice Problems and Additional Problems by Section" found under course materials on blackboard.

Calculators:

Students are expected to have a scientific calculator available during class every day, including days with quizzes or exams. The calculator may be a graphing calculator, or just a regular scientific calculator. Students may NOT use cell phones in place of calculators during quizzes or exams.

Attendance Policy:

Roll will be automatically recorded by collaborate. Although there are no direct points for attendance, the Top Hat questions can only be counted if you do attend class. Missing class will lower your attendance grade significantly, so you must attend class if you expect to pass this course.

Blackboard Collaborate

All lecture and lab classes will be held live in the scheduled time period. Please see the link for "Orientation for Collaborate" in the left column of your Blackboard Course.

Classroom Performance System (Top Hat):

 All students are required to enroll in the Top Hat classroom performance system at the beginning of the semester. You can use the app on your cell phone, tablet or computer or you can text your answers to the Top Hat phone number. Students who have not signed up by Tuesday, Jan. 25 at 8 am will not get credit for Top Hat activities until they do sign up. We will use the Top Hat application to respond to groups of multiple choice questions designed to enhance the learning experience in the course. Students will receive a participation grade based on the Top Hat questions. Students will be allotted Top Hat points based on simply answering the question, regardless of whether the responses are correct. Missing class will lower your Top Hat grade significantly, so you must attend class if you expect to pass this course. Top Hat points will be checked against attendance.

Achieve Learning Homework System:

Students will be provided with a list of assigned homework problems in the textbook from Chapters 1-4, 6, Sections 7.2-7.6 and Sections 8.1-8.4. There are two assignments for most, but not all of the chapters. These problems will be found in the Achieve Learning Online Homework System. Students will be graded on the completeness of the assignments. There will be no deduction for multiple attempts at problems. Students are expected to keep working all the homework problems until they have arrived at the correct answer. Do not hit "give up and view solution"! Students are expected to begin working on the homework for a particular problem set as soon as we begin talking about that material in the lab (first) and lecture (second). Due dates for homework assignments are listed in the chart below. Important: in order to get credit for your Achieve homework assignments, you *must* register for Achieve using the same last name as the one for you in the ASU Computer system. See the instructions for enrolling in Achieve Learning under content on your BlackBoard course.

Late Homework Policy:

The Achieve Learning Homework system can accept late homework assignments. However, for each day that the homework is late, 25% of the score will be deducted. After 4 days, no credit will be given for late assignments. The system computes the score if any deductions need to be made for late homework. Note that Achieve Homework 8-1, the last assignment, will not accept late assignments at all.

Policy on Academic Dishonesty:

Students are expected to work independently on quizzes, exams and lab reports. See the ASU Student Handbook for definitions of cheating and plagiarism. Any student who is caught cheating or plagiarizing in this class will be subject to failure in the course and possible suspension from the University. Cheating and/or plagiarism will not be tolerated! Angelo State University expects its students to maintain complete honesty and integrity in their academic pursuits. Students are responsible for understanding the Academic Honor Code, which is contained in both print and web versions of the Student Handbook.

Make-up Exam Policy:

Students will only be allowed to take one make-up exam if there is an illness or emergency which is documented in writing. In order for a student to be eligible to take a make-up exam, the student must notify Dr. Maxwell before the exam is missed by telephone, voice mail or email.

When a student returns to class after missing an exam, he or she <u>must</u> present a document such as a doctor's note or funeral notice in order for Dr. Maxwell to give permission for a make-up exam. All make-up exams will be given *on Friday, May 6 at 1:00 pm.* Students will not be given a second chance for a make-up exam. Each student may take no more than one make-up exam for any reason whatsoever.

Lab Course:

Attendance in lab is a mandatory part of this course. We will do group activities in small groups of 3 or 4 students. The point values for these activities can be found on the lab schedule on the last page of this syllabus. *Note that almost all of the new material will be presented during the lab period and not in the lecture. Poor attendance in lab will be fatal to your grade!*

Lab Reports:

Labs and Experiments will be carried out as group activities in lab. Each lab will be worth 30 points. Students are required to check their lab answers with the Instructor or Lab Assistant before leaving the lab that day. *The lowest lab report will be dropped, but if you miss lab with a documented excuse, you must carry out the work in the lab at home since quizzes will be given over the lab material near the beginning of the following week. Note that you are directly responsible for learning all the material from the lab each week.* No make-up labs will be accepted for any reason since your lowest lab grade will automatically be dropped. Only students with documented excuses are allowed to drop one lab.

Missing Lab Policy:

 Students will only be allowed to miss a lab if there is an illness, emergency or other schedule problem which is documented in writing. In order for a student to be eligible for a make-up lab, the student must notify Dr. Maxwell or the lab instructor <u>before</u> the lab is missed by telephone, voice mail or email. When a student returns to class after missing a lab, he or she <u>must</u> present a document such as a doctor's note or funeral notice or coach's note in order for the instructor to give permission to miss the lab. If you have a documented excuse, your missing lab grade will be dropped.   Note: Students can only be excused from one lab grade per semester.

Meets Tues/Thurs at 8:00-9:15 am in your Blackboard Collaborate Platform.

Week	Dates: Class Meetings the week of:	Sections Covered	Topics	Lab
1	**1/18** (Your Meetings are on TR)	**Sections 1.1 - 1.3**	**Pure substances, mixtures, periodic table, metals, nonmetals, metalloids,**	*Lab 2: Properties, Units and Conversions (25 pts)*
Saturday–Sunday, Jan 22-23	Achieve HW 1-1 due Saturday at 11:55 pm (8 pts)	Complete HW1-1 after reading through Lab 1 *Start Working on the HW the day you have lab 2.*	**Quiz 1 available from 10 am – 11:59 pm Sunday (will cover Lab 1). Don't forget we will do Lab 1 together in lecture class the first week.**	
2	**1/24** (Your Meetings are on TR)	**Sections 1.4 - 1.5**	**Metric units and prefixes, metric to metric conversions, significant figures, scientific notation, dosage calculations**	*Lab 3: Multistep Conversions, Dosage Calculations and Specific Heat (25 pts)*
Saturday–Sunday, Jan 29-30	Achieve HW 1-2 due Saturday at 11:55 pm (8 pts)	Complete HW1-2 after Lab 2 *Start Working on the HW the day you have lab 2.*	**Quiz 2 available from 10 am – 11:59 pm Sunday**	
3	**1/31** (Your Meetings are on TR)	**Section 1.6**	**Precision vs. accuracy, percent calculations in health science, physical and chemical changes, chemical reactions**	*Lab 4: Atoms, Isotopes and Radiation (25 pts)*
Saturday–Sunday, Feb 5-6	Achieve HW 1-3 due Saturday at 11:55 pm (9 pts)	Complete HW1-3 after Lab 3 *Start Working on the HW the day you have lab 3.*	**Quiz 3 available from 10 am – 11:59 pm Sunday**	
4	**2/7** (Your Meetings are on TR)	**Chapter 2**	**Atoms, subatomic particles, atomic number, mass number, isotopes, atomic mass Radioactivity, radioisotopes, nuclear equations, radioactive decay, radiations units and half-lives, medical applications for radioactivity**	*Lab 5: Arrangement of e-, Ionic Compounds & Introduction to Moles & Covalent Compounds (25 pts)*
Saturday–Sunday, Feb 12-13	Achieve HW 2 due Saturday at 11:55 pm (9 pts)	Complete HW2 after Lab 4 *Start Working on the HW the day you have lab 4.*	**Quiz 4 available from 10 am – 11:59 pm Sunday**	

Week	Dates: Class Meetings the week of:	Sections Covered	Topics	Lab
5	2/14 (Your Meetings are on TR)	Feb 17– Exam I (Chs 1 & 2) Administered during Class Time	Review for Exam 1	Review for Exam 1 (25 pts)
Saturday–Sunday, Feb 19–20		No Homework	No Quiz	
6	2/21 (Your Meetings are on TR)	Sections 3.1 – 3.3 & parts of 3.5	Electron arrangements, the octet rule, ion formation, ionic compounds, naming ionic compounds and writing the formula of ionic compounds with representative element ions, transition element ions and polyatomic ions, the mole, Avogadro's number, atoms to moles conversions, molar masses of compounds	Lab 6: Covalent Compounds and Moles (25 pts)
Saturday–Sunday, Feb 26–27	Ahieve HW 3–1 due Saturday at 11:55 pm (9 pts)	Achieve HW3–1 due Thursday at 11:55 pm (9 pts) *Start Working on the HW the day you have lab 5.*	Quiz 5 available from 10 am – 11:59 pm Sunday	
7	2/28 (Your Meetings are on TR)	Sections 3.4 – 3.7	Covalent bonds, electron dot structures, Lewis structures, grams to moles conversions, grams to atoms or molecules conversions, molar masses of compounds, grams to molecules conversions, Molecular geometry, electronegativity, polarity of bonds and molecules	Lab 7: Introduction to Organic Compounds and Families of Organic Compounds (25 pts)
Saturday–Sunday, March 5–6	Achieve HW 3–2 due Saturday at 11:55 pm (8 pts)	Complete HW3–2 after Lab 6 *Start Working on the HW the day you have lab 6.*	Quiz 6 available from 10 am – 11:59 pm Sunday	
8	3/7 (Your Meetings are on TR)	Sections 4.1 - 4.3	Alkanes, skeletal structures, functional groups, families of organic compounds, fatty acids	Lab 8: Naming Branched Alkanes and Isomers (25 pts)
		No Homework	No Quiz	
	3/14		SPRING BREAK	
		No Homework	No Quiz	
9	3/21 (Your Meetings are on TR)	Sections 4.4 - 4.5	Nomenclature of alkanes, types of isomers in organic chemistry, unsaturated fatty acids, chiral molecules in biochemistry	Lab 10: Intro to Carbohydrates: Monosaccharides (25 pts)

Week	Dates: Class Meetings the week of:	Sections Covered	Topics	Lab
	Achieve HW 4-1 due Saturday at 11:55 pm (8 pts)	Complete HW4-1 after Lab 7 *Start Working on the HW the day you have lab 7.*	**Quiz 7 available from 10 am – 11:59 pm Sunday**	
Saturday–Sunday, March 26-27				
10	3/28 (Your Meetings are on TR)	**Sections 6.1 - 6.4**	**Classes of carbohydrates, aldoses vs. ketoses, structures of common monosaccharides, stereochemistry of monosaccharides, cyclic forms of monosaccharides, oxidation and reduction of monosaccharides**	*Lab 11: Glycosidic Linkage, Disaccharides and Polysaccharides (25 pts)*
Saturday–Sunday, April 2-3	Achieve HW 4-2 due Saturday at 11:55 pm (8 pts)	Complete HW4-2 after Lab 8 *Start Working on the HW the day you have lab 8.*	**Quiz 8 available from 10 am – 11:59 pm Sunday**	
11	4/4 (Your Meetings are on TR)	**April 8– Exam 2 (Chs 3 & 4) Administered during Class Time**	**Review for Exam 2**	*Review for Exam 2 (25 pts)*
Saturday–Sunday, April 9-10		No Homework	**No Quiz**	
12	4/11 (Your Meetings are on TR)	**Sections 6.5 – 6.7**	**Hydrolysis of disaccharides, naming glycosidic linkages, labeling anomeric, acetal and hemiacetal carbon atoms, structures of common disaccharides, reducing disaccharides, structures of storage and structural polysaccharides, ABO blood types**	*Sections 2 & 3 from Lab 12 Lab 13: Attractive Forces, Solubility, Dietary Lipids and Cell Membranes (25 pts)*
Saturday–Sunday, April 16-17	Achieve HW 6-1 due Saturday at 11:55 pm (8 pts)	Complete HW6-1 after Lab 10 *Start Working on the HW the day you have lab 10.*	**Quiz 9 available from 10 am – 11:59 pm Sunday**	
13	4/18 (Your Meetings are on TR)	**Sections 7.2– 7.6 (Skip Section 7.1)**	**Polarity of organic molecules, types of attractive forces, liquids, solids, solubility, amphipathic compounds, soap, dietary lipids, trans fats, phospholipids, cellular membranes**	*Lab 14: Solutions, Solubility and Concentration (25 pts)*
Saturday–Sunday, April 23–24	Achieve HW 6-2 due Saturday at 11:55 pm (8 pts)	Complete HW6-2 after Lab 11 *Start Working on the HW the day you have lab 11.*	**Quiz 10 available from 10 am – 11:59 pm Sunday**	
14	4/25 (Your Meetings are on TR)	**Sections 8.1-8.4**	**Solutions, Solutes, Solvents, Colloids, Suspensions, Saturated and Unsaturated Solutions, Factors affecting Solubility, Chemical Equations for Solution Formation, Types of Electrolytes and Concentration**	*Review for Final Exam (25 pts)*

Week	Dates: Class Meetings the week of:	Sections Covered	Topics	Lab
Saturday–Sunday, April 30–May 1	Achieve HW 7-2 due Saturday at 11:55 pm (8 pts)	Complete HW7-2 after Lab 13 *Start Working on the HW the day you have lab 13.*	Quiz 11 available from 10 am – 11:59 pm Sunday	
15	5/2 (Your Meetings are on TR)		Review for Final Exam	No Lab
Saturday–May 7	Achieve HW 8-1 due Saturday at 11:55 pm (9 pts)	Complete HW8-1 after Lab 14 *Start Working on the HW the day you have lab 14.*	Note: no late assignments will be accepted for HW 8-1	
Final	Tuesday. May 10		Comprehensive Final Exam, including Chapters 7-8	Final Exam Administered 8:00 am – 10:00 am

The final exam for Section D10 will be held on Tuesday, May 10 at 8:00 – 10:00 am (available at 7:30 am)

Happy Summer!